

7.1 Compound Explanations

How to write Chemical Formulas and Chemical Names

Chemical Formula to Chemical Name

- Check to see if the first element in the compound is a Metal, Nonmetal or starts with H
If it is an "H"
Are there 2 elements or more than 2 elements
2 elements, Hydro ____ ic Acid
More, ate = ic, ite = ous Acid
If it is a Nonmetal...
The Subscript becomes the Prefix
(There is no prefix for the 1st element if there is only one)
If it is a Metal...
Check to see if it is a Transition Metal
If it is a Transition Metal, make the Table
If it is NOT a Transition Metal, write the name
(Double check for polyatomic ions)

Chemical Name to Chemical Formula

- Check to see if the first element in the compound is a Metal, Nonmetal or an Acid
If it is an Acid
Starts with Hydro = H and element
Does not start with Hydro
Ous = ite, ic = ate, H + polyatomic ion (cross and drop)
If it is a Nonmetal...
The Prefix become the Subscript
(There is no prefix for the 1st element if there is only one)
If it is a Metal...
Write out the chemical symbol with the charge of the metal/polyatomic ion
Write out the chemical symbol with the charge of the nonmetal/ polyatomic ion
Cross and drop
(Make sure to put polyatomic ions in () if there is more than one)

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Test Taking

Step 1 – Write down (Next to the problem #) if it is a metal (M), nonmetal (N), or acid (A)

Step 2 – Underline the endings (ide = PT, ate, ite = Sheet)

Step 3 – If it is a transition metal, Circle it

Step 4 – REPEAT for next problem

Step 5 – Complete ALL of one type of problem before moving on to the next type
i.e. complete all (A), then all, (N), etc.

Ex. M 1. FeO _____ Iron (II) Oxide _____

M 2. Sodium Phosphate _____ $\text{Na}_3(\text{PO}_4)$ _____

N 3. Carbon dioxide _____ CO_2 _____

A 4. Sulfuric Acid _____ H_2SO_4 _____